

showcasing research from Professor Sen Lin, School of Chemistry, Fuzhou University

selective hydrogenation of  $\alpha,\beta$ -unsaturated ketones by a single atom anchored on graphene: a study of dynamics

Based on density functional theory calculations, a reaction mechanism is proposed for the selective hydrogenation of  $\alpha,\beta$ -unsaturated ketones by a single atom anchored on graphene. The mechanism involves the coordination of the carbonyl oxygen to the Pd catalyst, followed by the migration of the Pd catalyst to the  $\beta$ -carbon of the double bond. The reaction proceeds via a transition state where the Pd catalyst is coordinated to both the carbonyl oxygen and the  $\beta$ -carbon. The final product is a saturated ketone. The study highlights the importance of the reaction mechanism in heterogeneous catalysis and offers an effective design principle for catalyzing heterogeneous chemical transformations.



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d a c a d a a ca ad  
 c . T c d c a b a  
 d a c c b ac  
 d a .  
 T a ad d a d c , ca  
 a . O P (111) a d Pd(111) ac ,  
 a , d c a (DFT) ca c a b  
 Sa a d c a a d a c ca b  
 a b d d ff c ad b 1,3-  
 b ad ( ad d C=C d b b d )  
 a d b ( a ad d C=C d b b d )  
 a a d ff ab a ad ca  
 d a .<sup>14</sup> H , c a c a  
 a SAC b ca ad  
 ac a d a b c d ff , ac  
 ac SAC ca b d a C=C d b b d .  
 L a d c c a d c  
 1,3-b ad ca a d b a A SAC d b a Z O<sub>2</sub>  
 ac .<sup>18</sup> T DFT ca c a a d a fi d -  
 a a OH d ac d  
 b ab ac H ad b d A , b  
 d a a d b a a d a a a  
 ba a fi d a . A  
 a a a d a ff d , b d c  
 d c a d .<sup>18</sup> W d  
 c ca a c c a a SAC d  
 , d b a cab SAC a <sup>11</sup>  
 2D a a ,<sup>12</sup> a OH a b  
 a .  
 I , b c b a c DFT ca c a ab  
 initio c a d a c (AIMD)<sup>19</sup> d c d a c <sup>20</sup>  
 a , 1,3-b ad d a  
 a a d Pd SAC (Pd<sub>1</sub>/ a ) a a d .  
 T a ff a c a d -  
 a d c b ad d a a c -  
 a -d fi d SAC.<sup>12</sup> O DFT ca c a a a c a  
 a ac - c SAC a d a  
 ca a c a a , c a d b c  
 a b a . P a ,  
 c a d b AIMD a b  
 - a - a d a c c d d a  
 , c 1-b d b a c a  
 a d a Pd ffic c  
 c ad . T d a ca a  
 a a a a b d  
 a C=C d b b d 1-b Pd, c  
 d ad a d a b a a d b  
 c a b a .  
 T a c d a c ca a a  
 c a ac d c a . F a da a  
 d ca c , d a d a  
 a d ff ac a ( b a vs. a -  
 a ) d ca d ff fficac .<sup>21-23</sup>  
 H , a b a fl c  
 d a c d c b a c ac c . T  
 d a ca d c d a d c c -  
 a a dd c a c c a

ac (PES) d c a c a a d  
 ac , b a d c b a c . S c fica  
 ca d a ca a ad ,<sup>24</sup> a ac d  
 c ac c d a (IRC) PES, c d fi  
 a ac a a a  
 a d a d c . H , c a  
 a a d ca d c d ca a  
 ca ba d c a b a d a c ca  
 ac , c a da c IRC aft a  
 ac a a ad “ c d”  
 d c .<sup>25-31</sup> T ac -IRC ac c  
 , c acc d b PES a . I a  
 d a c c a ac a ac , a a , c  
 a d aft a a a ca  
 ad a c a a . I  
 d c d , d a ca d c  
 d a 1,3-b ad d a a a  
 a a d ac  
 ca a . U d c c c a c , d a c a  
 b c c d d .

## Computational details

### DFT

A - a d d c a (DFT) ca c a  
 ca d V a *Ab initio* S a Pac a  
 (VASP)<sup>32,33</sup> ad -c c d P d -B -E  
 (PBE) c a .<sup>34</sup> T a c c d c b d  
 a a - a ba a c ff 400 V,  
 c c a a d c  
 a d- a d .<sup>35</sup> T a c c d  
 d ac b G - ca  
 c <sup>36</sup> a d a d Waa  
 c c a d d Tab S1.† A b c  
 c a a = 9.84 Å ( a x d c -  
 ), b = 12.78 Å (fi a y d c ) a d c = 30.0  
 Å d . A 2 × 2 × 1 M -Pac k- d<sup>37</sup>  
 ad d B a a d c c .  
 T d b c a ad  
 a a c a a a  
 0.03 V Å<sup>-1</sup> , c a a ca a a d ad ba  
 a d . T c b a d d a c ba d (CI-  
 NEB)<sup>38,39</sup> a ac a d d ac  
 a a d ba a a a d  
 a a c a ac c ca a  
 a d a a .

### AIMD simulations

T - a - a d a c , *ab initio* c a  
 d a c (AIMD) a <sup>19</sup> ca d VASP  
 a a d c b d ab . A c a ac c  
 a c fi a a d b a a a a  
 d d b NEB ca c a a c c .  
 F a c c d a a  
 a a a , b a c  
 a a d a d a d ca d a a



300 K. W a d c ca ca  
 a a c d a c  
 a d a a a ,<sup>40,41</sup> c a b a d  
 d a a a - d H a a d  
 a a . T c ac a a d a  
 300 K. T a c a a d a c -  
 ca ca (NVE) b a - a  
 d VASP, a a a a 2  
 a da 0.5 . T a a c d  
 ~20 V a a c .

## Results and discussion

### Model of the Pd<sub>1</sub>/graphene SAC

H a Pd SAC a fi  
 a d a c d a d ca a c c a  
 c d a 1,3-b ad . T d  
 d d ba d a  
 Ya *et al.*,<sup>12</sup> d ab ca Pd- a  
 a a X- a c c (XPS),  
 ca a c c c (STEM), a d  
 d d X- a ab fi c (EXAFS) a -  
 . T a da a a a ca  
 d d Pd c a a c d a a-  
 c d a C a d O a .  
 O d Pd/ a SAC c Pd-  
 C a d Pd-O b d , a F . 1, c  
 b d fi - a EXAFS c .<sup>12</sup> T  
 c O<sub>a</sub>, c b d Pd a a d a , c -  
 d a c c a d b da a .<sup>12</sup>  
 T Pd a a c d a C<sub>a</sub>. T  
 c a-c d a Pd, c  
 (O<sub>b</sub> a d O<sub>b'</sub>) a c d d (F . 1a), c a ab  
 d aft a afl ac ac a  
 a d d a c a d Pd c .<sup>12</sup>  
 F , d a a add a d d  
 a a ca b a a d c d  
 a d a add a Pd-C b d . T  
 C-O-C d d a  
 a d a a a c a a a ca a c . A  
 Tab 1, d SAC d  
 b d a a d a a

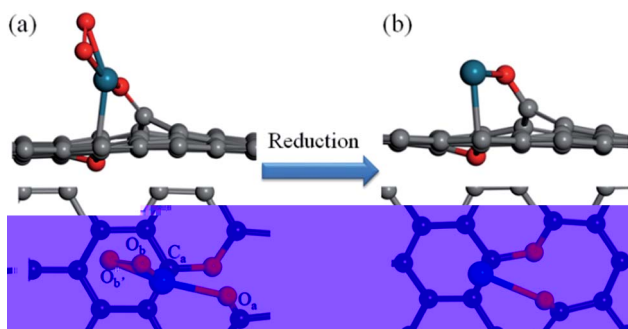
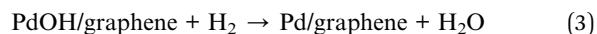
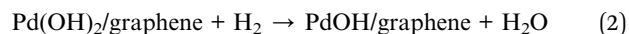
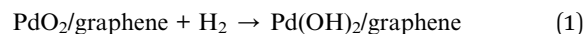


Fig. 1 Schematic and 3D model of Pd<sub>1</sub>/graphene SAC (a) schematic and (b) 3D model. C, grey; O, red; Pd, blue.

Table 1 Calculated adsorption energies of H<sub>2</sub>, O<sub>2</sub>, and CO<sub>2</sub> on Pd SAC

Adsorbate	Adsorption energy (eV)	Reference
H <sub>2</sub>	2.09	2.05
O <sub>2</sub>	2.05	2.07
CO <sub>2</sub>	2.00	2.07
Pd-C <sub>a</sub>	2.25	2.00

Pd-C<sub>a</sub> d Pd-O<sub>b</sub> d fi c d -  
 .<sup>12</sup> I add , c c d ca b  
 a d a c a a  
 EXAFS.<sup>12</sup> T Bad c a a a d ca a c a  
 Pd a +0.69e, c a b -  
 a a a d Pd a a ac a  
 c a d.<sup>12</sup>  
 T SAC c d c d ab a  
 a ac ca a ba d a DFT ca c a  
 d a ac ( ), ab  
 b ca c ca a-c d a d Pd. I  
 d a ca a a -  
 a d 10% H<sub>2</sub> A a 150 °C, c c d c c ab  
 d c O b d d Pd. T b ,  
 ac (1)-(3) a a d ca :



I a d DFT ca c a a ac (1) a  
 a a c -3.28 V, d ca a Pd  
 b d O ca b ad d c d d (OH)  
 c H<sub>2</sub> (F . S1 ESI<sup>†</sup>). T a d  
 ca b d c d b a add a H<sub>2</sub> c  
 H<sub>2</sub>O c ac (2) a d (3). T  
 ac a ca c a d a d a ba , 0.98 a d  
 0.63 V, c (F . S2<sup>†</sup>). T fi a ac a d SAC  
 a a a d Pd a a c d a  
 (F . 1b).

T Pd/ a SAC d d b  
 a b ad a d ca a . T SAC  
 a ab ab , a Pd b d -1.87 V c  
 ca c a d E<sub>b</sub> = E<sub>Pd-</sub> - E<sub>ads</sub> - E<sub>Pd</sub> E<sub>Pd-</sub>  
 Pd-SAC, E  
 b Pd a a d E<sub>Pd</sub> Pd  
 a . T a Pd ad (-1.75 V)  
 C O<sub>2</sub>(111).<sup>42</sup> T Pd a a c a +0.30e,  
 a fica ca c c a ac d c a a  
 a d .

### Adsorption of pertinent species

B a ac c a , ad  
 b a H<sub>2</sub>, 1,3-b ad (13BD), 1-b (1B), a d b a  
 d d (F . 2). T d c d ad b



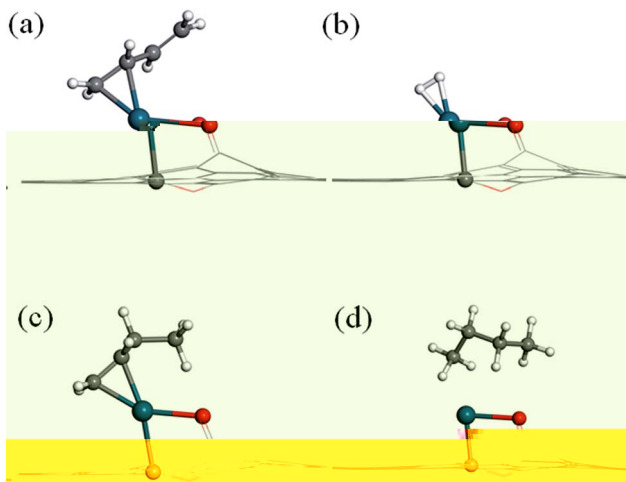


Fig. 2 Adsorption of (a) 1,3-butadiene, (b) H<sub>2</sub>, (c) 1-butene and (d) butane on the Pd(111) surface. C (grey), O (red), Pd (blue), H (white). A black arrow indicates the direction of the reaction.

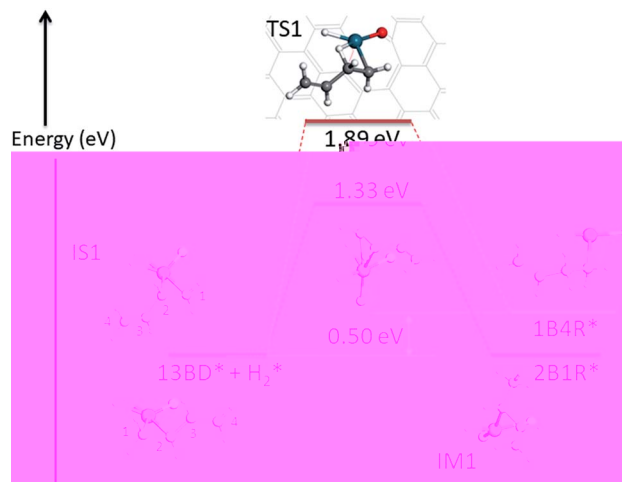


Fig. 3 Energy profile of the reaction from the initial state (IS1) to the final state (FS) via the transition state (TS1). IS1: initial state; TS1: transition state; IM1: intermediate state; FS: final state. The energy profile is shown in eV.

Pd–H distance 1.73 Å and 1.74 Å, respectively, and the Pd–C distance is 2.13 Å and 2.17 Å, respectively. The barrier height is 1.66 eV, which is very close to the experimental value of 1.65 eV.

For 1,3-butadiene, the adsorption energy is 1.33 eV. The transition state energy is 1.89 eV, and the intermediate state energy is 1.57 eV. The final state energy is 1.24 eV.

For 1-butene, the adsorption energy is 1.09 eV. The transition state energy is 1.89 eV, and the intermediate state energy is 1.57 eV. The final state energy is 1.24 eV.

For butane, the adsorption energy is 0.87 eV. The transition state energy is 1.89 eV, and the intermediate state energy is 1.57 eV. The final state energy is 1.24 eV.

Reaction mechanism of 1,3-butadiene hydrogenation

The reaction mechanism involves the adsorption of 1,3-butadiene on the Pd(111) surface. The initial state (IS1) is at 0.50 eV, and the transition state (TS1) is at 1.89 eV. The intermediate state (IM1) is at 1.57 eV, and the final state (FS) is at 1.24 eV. The reaction energy is 0.81 eV.

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T b c c d d a  
 a a . O a d, 1-b  
 d c ca c Pd, a F .4, a  
 π-ad C<sub>3</sub>=C<sub>4</sub> (ad  
 -1.72 V). I F .S3,† a c c  
 a a (TS2) a d ad b d 1-b (1B\*) a d  
 c a a fica a  
 c . T ca a d a c d d  
 d a b a , c  
 a b a . O a d, 1-b  
 d c ca d b, ad a d c a a  
 -0.65 V, a a fi . I  
 d ac , d d a , b d  
 b c a b d c-  
 a d 1-b , a b a d a  
 C<sub>3</sub>=C<sub>4</sub> d b b d d b ad ca a d b a Pd  
 SAC.

### Post-transition-state dynamics

B c TS2, F .4, d ffic a  
 C<sub>3</sub>=C<sub>4</sub> a Pd a d c 1-b  
 c d d d d b, a a IRC a d  
 ad b d 1B\* F .S3,† T d c  
 d c a , fi d a AIMD ca c a-  
 a TS2 (F .5a) a c a  
 a . T fi a (2 ) a a c  
 F .5b, c a 1-b d c a  
 d ab. 15 Å a a Pd. T  
 a c ca b d SI. T c fi fi d ,  
 add a a c a d a a TS2  
 , b a d a d a c a d 16  
 a d a c d a  
 ac b a d ff d ac d a  
 a d a ac c d a . A  
 a c ad d 1-b ,  
 ad 1-b Pd SAC.



a c<sup>25-30</sup> a d a c ac ,<sup>31</sup> a  
a b d c SAC-ca a d ac  
I , c a c da a  
d ff a P (111) a d Pd(111), d ff  
ad 1,3-b ad a d 1-b c b -  
a d ff ab a ad ca  
d a a b c d c 1-  
b .<sup>14</sup> T , d a c d d a a  
c .

### Conclusions

I , add a  
d a a c d a 1,3-b ad  
ca a d b a Pd SAC a c d a . F , DFT  
ca c a d c da a SAC ac  
a d ca a c c a . O Pd SAC d  
c a c a a c a a . F  
c a ca H -P a c a , d a  
1,3-b ad 1-b d a a  
c d a d add C<sub>1</sub>=C<sub>2</sub>  
d b b d Pd-ad b d 1,3-b ad b Pd ac a d  
H<sub>2</sub>. P a a , d a ca d 1-b  
d aft c d d a a d -  
ad d c Pd a d b  
b da d a d a . T ca  
c d d ca a a d d c  
a c d a c . T c a c  
a b a SAC c a d ff a ff c -  
d c SAC ca a d c -  
ca a a .

### Conflicts of interest

T a c flc d ca .

### Acknowledgements

W ac d Na a Na a Sc c F -  
da C a (21673040 S. L.), Na a Sc c F da  
F a P c (2016J01052 S. L.), F U  
Q a Sc a P a (XRC-17055 S. L.), a d U.S.  
Na a Sc c F da (CHE-1462109 H. G.). L. Z.  
ac d a a C Sc a  
C c (201604890009).

### Notes and references

- 1 M. L. D , *Stud. Surf. Sci. Catal.*, d. L. C ,  
E , 1986, . 27, . 613-666.
- 2 T. O c a b, J. Ma a d a d A. R , *J. Catal.*, 1989,  
119, 517-520.
- 3 A. Sa a , Z. Z d , G. S fl , J. W. H a d  
L. G c , *J. Catal.*, 1995, 157, 179-189.
- 4 J. G , M. A. V , C. E. G a a d R. T . d , *J. Catal.*,  
2001, 199, 338-345.

- K. N. Houk, *Proc. Natl. Acad. Sci. U. S. A.*, 2018, **115**, E848–E855.
- 32 G. K. Kulkarni and J. F. Liebman, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1996, **54**, 11169–11186.
- 33 G. K. Kulkarni and J. F. Liebman, *Comput. Mater. Sci.*, 1996, **6**, 15–50.
- 34 J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, **77**, 3865–3868.
- 35 P. E. Blöchl, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1994, **50**, 17953–17979.
- 36 S. Grimme, J. Antonysson, S. Ehrlich and H. Klöpper, *J. Chem. Phys.*, 2010, **132**, 154104.
- 37 H. J. Morosini and J. D. Paciorek, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1976, **13**, 5188–5192.
- 38 G. Henkelman, B. P. Uberuoba and H. Jónsson, *J. Chem. Phys.*, 2000, **113**, 9901–9904.
- 39 G. Henkelman and H. Jónsson, *J. Chem. Phys.*, 2000, **113**, 9978–9985.
- 40 S. Carles and D. L. Bredas, *J. Chem. Phys.*, 1975, **62**, 2890–2899.
- 41 W. L. Hase and D. G. Bredas, *Chem. Phys. Lett.*, 1980, **74**, 284–287.
- 42 Z. Yan, Z. Li, G. Li and K. Han, *Phys. Lett. A*, 2007, **369**, 132–139.
- 43 G. Garcia-Castano, F. J. Cadogan-Saiz and J. C. Bozzano, *J. Catal.*, 2003, **214**, 26–32.

